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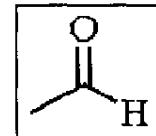
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Acetaldehyde



[75-07-0] · C₂H₄O · Acetaldehyde · (MW 44.05)

(reagent used as two-carbon electrophilic component in a wide array

Physical Data: mp -123.5 °C; bp 21 °C; d 0.788 g cm⁻³.

Solubility: sol H₂O, alcohol, ether, and most organic solvents.

Form Supplied in: colorless liquid; widely available.

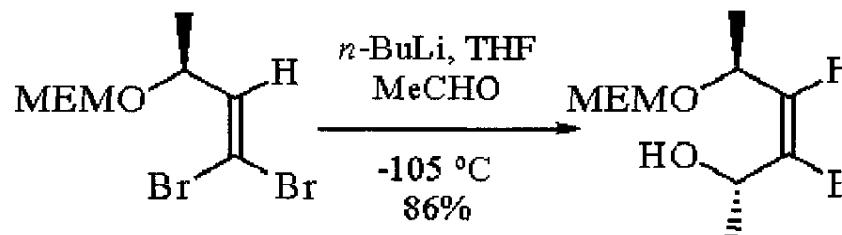
Purification: shaken with powdered NaHCO₃ for 30 min; dried over 760 mmHg through a 70 mm Vigreux column.

Handling, Storage, and Precautions: bottles may develop pressure as To help prevent polymerization and autoxidation, store under nitrogen. Acetaldehyde is a cancer suspect agent and should be used only in a fume hood. (oral) rat LD₅₀: 661 mg kg⁻¹. Incompatible with strong acids, strong bases, and strong oxidizing agents. Decomposes on prolonged exposure to air.

1,2-Additions.

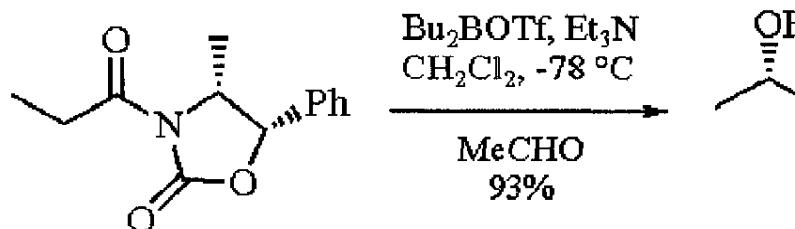
Acetaldehyde reacts with a myriad of nucleophilic reagents, generally two-carbon extended secondary alcohols. Aryl-,¹ alkynyl-,² and alkyne acetaldehyde even at low temperature. A chiral vinylolithium reagent reacts stereoselectively to afford a 10:1 mixture of diastereomeric alcohols.³ Reagents behave in an analogous manner with acetaldehyde to give the corresponding ketones⁷ upon subsequent oxidation. Allyl organometallics react with

depending on the metal and conditions to give the corresponding homallylboronates also react with acetaldehyde at -78 °C to afford the homoenantioselectivity.⁹ *trans*-Epoxides are produced selectively through acetaldehyde with halomethyl sulfones under basic phase transfer conditions¹¹ and Horner-Emmons phosphonate¹² ylides react with acet



Aldol Additions.

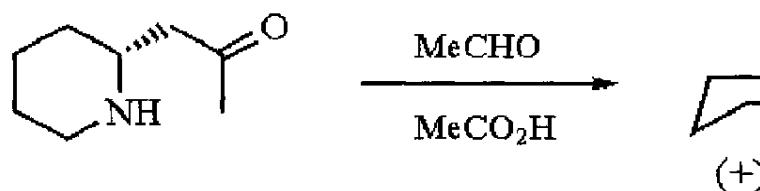
Acetaldehyde serves as an electrophilic partner in the aldol condensation. Knoevenagel condensation of acetaldehyde with active methylene compounds provides good yields of the ethylenic substituted compounds.¹⁴ Addition of a methylene compound to acetaldehyde results in a Michael addition or in the initially formed ethylenic compound.¹⁵ Tollens reaction of acetaldehyde with tetahydrofuran yields pentaerythritol.¹⁶ The addition of acetaldehyde in a Baylis-Hillman reaction using *1,4-Diazabicyclo[2.2.2]octane* (DABCO) as catalyst gives a 96% stereoselective aldol reaction of acetaldehyde with achiral¹⁸ and chiral¹⁹ aldehydes. Much attention and is a proven method for controlling acyclic relative stereochemistry. For example, the boron enolate of a norephedrine-derived propionylcyclohexene undergoes nitro-aldol condensation to the corresponding nitro alcohol.²⁰ A variety of heterocycles react with acetaldehyde to give good yields of the product.²¹ Zinc,²² copper,²³ and boron²⁴ enolates of esters and ketones provide



Mannich and Mannich-Type Reactions.

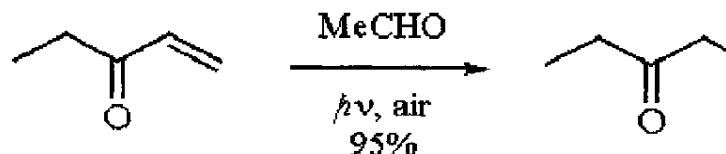
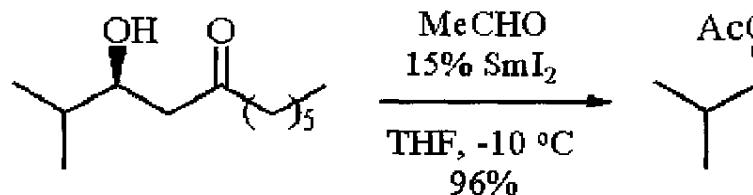
Although not as commonly used as *Formaldehyde*, acetaldehyde and

Mannich reactions. Intramolecular Mannich reaction of acetaldehyde natural product myrtine (eq 3).²⁵ The intramolecular Mannich reaction synthesis of proline derivatives.²⁶ Nucleophiles as diverse as dialkyl radicals²⁹ may also add to the intermediate imine of acetaldehyde in historically significant reaction of acetaldehyde in this mode is the S_N cyanide is added to the adduct of ammonia and acetaldehyde followed by α -aminonitrile.³⁰ The Pictet-Spengler reaction utilizing acetaldehyde reaction. Acetaldehyde has been extensively used in the synthesis of from tryptophan derivatives through this cyclization.³¹ Other ring systems tetrahydroisoquinolines³² and dihydrooxazines³³ have also been formed by cyclization with acetaldehyde.



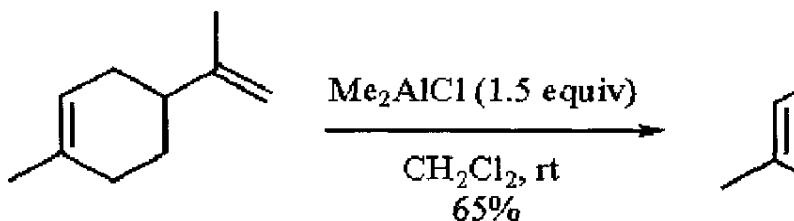
Metal and Other Promoted Condensations.

In the mixed Tishchenko reaction using *Aluminum Isopropoxide* as the predominately the oxidized partner. Thus when condensed with benzene, the major product.³⁴ Recently an interesting and synthetically useful stereoselective Tishchenko reduction of β -hydroxy ketones using acetaldehyde and aluminum isopropoxide, affording *anti*-1,3-diol monoacetates, has been reported (eq 4).³⁵ The coupling of acetaldehyde with other higher-order aldehydes that contain a carbonyl group has been achieved using a vanadium(II) reagent.³⁶ The photochemical addition of acetaldehyde in the presence of molecular oxygen to α,β -unsaturated esters and ketones to form cyclic dicarbonyl compounds (eq 5).³⁷



Pericyclic Reactions.

The thermal ene reactions of acetaldehyde and other aliphatic aldehydes are very productive.³⁸ However, acetaldehyde can be induced to undergo ene reactions under Lewis acid activation. **Dimethylaluminum Chloride** has been employed to obtain reasonable yields of ene products with acetaldehyde as the unreactive dieneophile towards dienes. The hetero-Diels-Alder reaction has been reported under high pressure acceleration with 1-alkoxydienes to afford ene products with modest *endo* selectivity.⁴¹



Paraldehyde and Other Acetaldehyde Derivatives

Paraldehyde has historically been used as a stable and less volatile form of chemical reactions.⁴² However, since its classification as a controlled substance has led to its limited use in modern synthetic organic chemistry, it is often generated from paraldehyde through acid catalyzed degradation of the diethyl acetal.⁴³ The diethyl acetal of acetaldehyde, commonly known as diethyl acetal, is formed by the reaction of acetaldehyde, ethanol, and calcium chloride.⁴⁴ Acetal has been used for the protection of diols as their ethyldene acetals.⁴⁵

Related Reagents.

Acetaldehyde N-t-Butylinime; Acetaldoxime; Crotonaldehyde; Dimethylvinyl Ether; Formaldehyde; Formaldehyde-Dimethylamine; Vinyl

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